

Unit 6C Molecular Formulas Practice II

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V	work each of the following problems. Show ALL WORK.				
1.	The empirical formula of a compound is found to be P_2O_5 . Experiments show that the molar mass of the compound is 283.9 g/mol. What is the molecular formula of the compound?				
2.	A compound has the following % composition — 76.54 % C, 12.13 % H, and 11.33 % O. If its molar mass is 282.5 g/mol, what is its molecular formula?				
3.	What is the formula for a hydrate which consists of 90.7 % $\mathrm{SrC_2O_4}$ and 9.30 % $\mathrm{H_2O?}$				