Lab: Redox Reactions—Datasheet **Name**_____

Complete and balance the equations for the single displacement reactions you observed. Then assign an oxidation number to each element.

$Cu + AgNO_3 \rightarrow$

Assume copper II nitrate is formed.

$Al + CuCl_2 \rightarrow$

$Mg + CuSO_4 \rightarrow$

Conclusions:

1. How can you tell that the reactions were all redox reactions?

2. When a metal goes into solution, it (***loses, gains***) electrons and is (***oxidized, reduced***) to become a (***+, –***) ion. At the same time, it forces the metallic ion already in solution to (***lose, gain***) electrons and come out of solution as an atom. The ion is (***oxidized, reduced***). This particular redox reaction is called single _____.

3. In each equation above, circle the reactant that is oxidized in the reaction and draw a square around the reactant that is reduced.