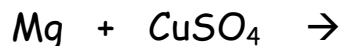


Complete and balance the equations for the single displacement reactions you observed. Then assign an oxidation number to each element.



Assume copper II nitrate is formed.



Conclusions:

1. How can you tell that the reactions were all redox reactions?
2. When a metal goes into solution, it (*loses, gains*) electrons and is (*oxidized, reduced*) to become a (+, -) ion. At the same time, it forces the metallic ion already in solution to (*lose, gain*) electrons and come out of solution as an atom. The ion is (*oxidized, reduced*). This particular redox reaction is called single _____.
3. In each equation above, circle the reactant that is oxidized in the reaction and draw a square around the reactant that is reduced.