Problem Set One: Episode 803

$$C_5H_{12} + O_2 \rightarrow CO_2 + H_2O$$

BALANCE THE EQUATION!!!!!!

Given that the density of carbon dioxide is approximately 1.99 g/L, what *volume*, in liters, of carbon dioxide will be produced if 85.0 g of pentane are burned?

How many *molecules* of water will be produced if 26.3 g of pentane are burned?