

electrochemistry - the study of \_\_\_\_\_-related applications of \_\_\_\_\_ - \_\_\_\_\_ reactions

oxidation numbers -

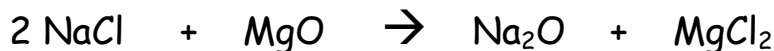
- numbers assigned to the \_\_\_\_\_ in a chemical \_\_\_\_\_ that give the \_\_\_\_\_ charge of the \_\_\_\_\_.
- In ionic compounds - oxidation # = \_\_\_\_\_
- In molecular compounds - oxidation # is based on \_\_\_\_\_
- Oxidation #'s are written with the \_\_\_\_\_ first and then the \_\_\_\_\_.

Rules For Assigning Oxidation Numbers:

situation	oxidation #

elements in a compound:


Example:



An \_\_\_\_\_-\_\_\_\_\_ reaction ( \_\_\_\_\_ reaction) is any reaction in which atoms or ions undergo a \_\_\_\_\_ in \_\_\_\_\_.

Is this a redox reaction?  $2 \text{Na} + \text{Cl}_2 \rightarrow 2 \text{NaCl}$  \_\_\_\_\_

When an \_\_\_\_\_ or \_\_\_\_\_ in a reaction has an \_\_\_\_\_ in \_\_\_\_\_ number, it has undergone the process of \_\_\_\_\_ by \_\_\_\_\_ 1 or more \_\_\_\_\_.

The oxidation # of Na went from \_\_\_\_\_ to \_\_\_\_\_, so Na was \_\_\_\_\_.

When an \_\_\_\_\_ or \_\_\_\_\_ in a reaction has a \_\_\_\_\_ in \_\_\_\_\_ number, it has undergone the process of \_\_\_\_\_ by \_\_\_\_\_ 1 or more \_\_\_\_\_.

The oxidation # of Cl went from \_\_\_\_\_ to \_\_\_\_\_, so Cl was \_\_\_\_\_.

LEO says GER

\_\_\_\_\_ electrons - \_\_\_\_\_      \_\_\_\_\_ electrons - \_\_\_\_\_

Chemistry Assignment #3: On Assignment #2, draw circles around the reactants that are oxidized and squares around those that are reduced.

### 3 Types of Redox Reactions:

#1 Electron exchange happens \_\_\_\_\_, as the chemicals come in \_\_\_\_\_.

#2 Electron exchange is forced, using \_\_\_\_\_. This is called \_\_\_\_\_.

#3 The chemicals are \_\_\_\_\_ so that \_\_\_\_\_ exchange occurs only when the electrons move through a \_\_\_\_\_, producing \_\_\_\_\_.

Chemistry Quiz:

1. \_\_\_\_\_ 2. \_\_\_\_\_ 3. \_\_\_\_\_ 4. \_\_\_\_\_ 5. \_\_\_\_\_  
CR 1 \_\_\_\_\_ CR 2 \_\_\_\_\_