

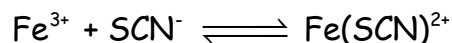
Table I:

Ion	Color
K ⁺	
Cl ⁻	
SCN ⁻	
Fe ³⁺	
Fe(SCN) ²⁺	

Use the chart on the left to determine the colors of reactants and products. Write the color in the blanks above the equation.

Table II:

colors: _____



	Chemical Added	Color Change	Direction of Shift
1.	FeCl ₃		
2.	NaOH		
3.	KSCN		

CONCLUSIONS:

According to LeChatelier's Principle, when a _____ is applied to a system in equilibrium, the system will readjust to _____ the stress, restoring a state of equilibrium.

For each procedure in Table II, identify the stress (ex.- addition of a reactant, removal of a product, etc.) and the reason for the shift in equilibrium (ex.- shift to right uses up reactants):

Stress

Reason for Shift

- _____
- _____

Hint: NaOH reacts with Fe³⁺ to form solid Fe(OH)₃.
- _____