

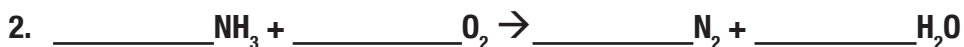
Answer each of the following questions using the equation provided. BE SURE TO BALANCE EACH EQUATION BEFORE SOLVING ANY PROBLEMS and SHOW ALL WORK.



a. 2 moles of NO will react with _____ mole(s) of O₂ to produce _____ mole(s) of NO₂.

b. ? moles NO₂ = 3.6 moles O₂ × $\frac{\text{_____ moles NO}_2}{\text{_____ moles O}_2}$ = _____

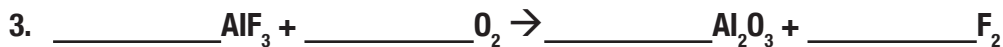
c. How many moles of NO must react to form 4.67 moles of NO₂?



a. 20 moles of NH₃ are needed to produce _____ moles of H₂O.

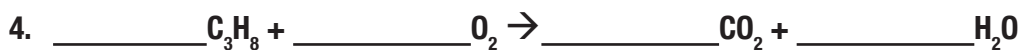
b. How many moles of N₂ will be produced if 3.5 moles of O₂ react?

Answer each of the following questions using the equation provided. **BE SURE TO BALANCE EACH EQUATION BEFORE SOLVING ANY PROBLEMS and SHOW ALL WORK.**



a. 20 moles of AlF_3 will produce _____ moles of F_2 .

b. _____ moles of AlF_3 will react with 0.6 moles of O_2 .



a. How many moles of oxygen react with 11 moles of C_3H_8 ?

b. How many moles of CO_2 are produced if 3.5 moles of water are produced?



a. Fill in the following word equation — _____ moles of oxygen gas react with _____ moles of iron to produce _____ moles of iron (III) oxide.

b. _____ moles of O_2 are required to produce 3.0 moles of iron (III) oxide.