CHEMISTRY MATTERS
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Unit 5E Practice Questions IV
Determining Oxidation Number

Name:

Date:

Complete and balance the equations for the single displacement reactions you observed.
Then assign an oxidation number to each element.

Assume copper II nitrate is formed

1. $Cu + AgNO_3 \rightarrow$

2. $Al + CuCl_2 \rightarrow$

3. $Mg + CuSO_4 \rightarrow$

Conclusions:

1. How can you tell that the reactions were all redox reactions?

2. Choose the best answer from words within parenthesis below and underline it.
   When a metal goes into solution, it (loses, gains) electrons and is (oxidized, reduced) to become
   a (+, -) ion. At the same time, it forces the metallic ion already in solution to (lose, gain) electrons
   and come out of solution as an atom. The ion is (oxidized, reduced). This particular redox reaction
   is called single _____.

3. In each equations above, circle the reactant that is oxidized in the reaction and draw a square around the
   reactant that is reduced.