CHEMISTRY MATTERS
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Unit 5E Balancing Equations Practice Problems I

Name:

Date:

A chemical equation tells the story of a chemical reaction. Reactants are the starting substances in the reaction while products are the new substances that are formed. The large numbers in front of some of the formulae are called coefficients. These coefficients are used to balance the equation. Chemical reactions must obey the Law of Conservation of Matter. The number of atoms of each element on both sides of the equation must be equal because matter cannot be created or destroyed. When balancing equations, the only numbers that can be changed are coefficients. Subscripts in a chemical formula cannot be changed to balance an equation.

Balance the following equations:

1.	Al	+	O_2	→	Al_2O_3			
2.	$_3H_8$	+	O_2	→	CO_2	+	H_2O	
3.	$Al(NO_3)_3$	+	$NaOH$	→	$Al(OH)_3$	+	$NaNO_3$	
4.	KNO_3	→	KNO_2	+	O_2			
5.	O_2	+	CS_2	→	CO_2	+	SO_2	
6.	$KClO_3$	→	KCl	+	O_2			
7.	BaF_2	+	K_3PO_4	→	$Ba_3(PO_4)_2$	+	KF	
8.	H_2SO_4	+	$Mg(NO_3)_2$	→	$MgSO_4$	+	HNO_3	
9.	Al	+	H_2SO_4	→	$Al_2(SO_4)_3$	+	H_2	
10.	WO_3	+	H_2	→	W	+	H_2O	