1. A student looked up the naturally occurring isotopes of bromine and found the following information:

50.54% of the naturally occurring isotopes of bromine have an atomic mass of 78.92 u while 49.46% of the naturally occurring isotopes of bromine have an atomic mass of 80.92 u.

Calculate the average atomic mass of bromine, showing all work:

2. Using the following data, calculate the average atomic mass of magnesium (give your answer to the nearest .01 u) : Show all work!

<table>
<thead>
<tr>
<th>Isotope</th>
<th>Percent abundance</th>
</tr>
</thead>
<tbody>
<tr>
<td>$^{24}\text{Mg}$</td>
<td>78.70%</td>
</tr>
<tr>
<td>$^{25}\text{Mg}$</td>
<td>10.13%</td>
</tr>
<tr>
<td>$^{26}\text{Mg}$</td>
<td>11.17%</td>
</tr>
</tbody>
</table>

3. Using the periodic table,

What is the average atomic mass of bromine? 

What is the average atomic mass of magnesium? 

How do your calculated answers in #1 and #2 compare to those on the periodic table