

Main Ideas, Key Points, Questions:

After watching the video segment, write down key points, main ideas and big questions.

NOTE-TAKING GUIDE UNIT 3, SEGMENT E

Name:

Date:

Objective(s):

To use atomic data to support an argument that explains changes in effective nuclear charge, ionization and atomic radius in the periodic table.

Notes:

During the video segment, use words, phrases or drawings to take notes.

Summary:

After watching the video segment, write at least three sentences explaining what you learned. You can ask yourself: "If I was going to explain this to someone else, what would I say?"



QUESTIONS TO CONSIDER: UNIT 3, SEGMENT E Name:

Date:

After watching the video and performing any associated labs and/or experiments, you should be able to answer the following:

- 1. When effective nuclear charge decreases, what happens to the attraction between the nucleus and the electrons?
- 2. What are core electrons?

- 3. Sodium is a larger atom than lithium. List three reasons why this is true.
 - 1. 2. 3.
- 4. Effective nuclear charge increases as you move to the right across periods on the periodic table. What causes the increase in effective nuclear charge?
- 5. What causes ionization to get stronger as you move from left to right across the periodic table?

6. What causes atoms to get smaller as you move from left to right across the periods?