

Fill in the blanks spots of the table below.

Ionic compounds are composed of cations and anions. When writing an ionic formula, the total positive charge must equal the total negative charge because the number of electrons lost must equal the number of electrons gained. Always write the symbol of the cation first, followed by the symbol of the anion. The criss-cross method is a short-cut to correct formula writing as long as we remember to simplify our subscripts!

In each box, write the formula of the ionic compound consisting of the positive ion to the left of the box and the negative ion above the box.

	Cl^-	S^{2-}	F^-	N^{3-}	O^{2-}	P^{3-}
Mg^{2+}						
Cs^+						
Cr^{3+}						
Na^+						
Zn^{2+}						
Al^{3+}						