1. Give the word equation for the neutralization reaction of an acid and a base.

2. Complete these equations:
   
   \[ \text{HCl} + \text{LiOH} \rightarrow \]
   
   \[ \text{HC}_2\text{H}_3\text{O}_2 + \text{Mg(OH)}_2 \rightarrow \]

3. A ____________________ is a laboratory method used to determine the concentration of an acid or a _______________ in solution by performing a ________________ reaction with a standard solution.

4. At the _______________ of the titration, the indicator changes color, which indicates neutralization. Once neutralized, moles of _______________ and moles of _______________ are equal.

5. In a titration of HCl with NaOH, 100.0 mL of the base was required to neutralize 20.0 mL of 5.0 M HCl. What is the molarity of the NaOH? (Be sure to write the neutralization reaction.)

6. In a titration of H\textsubscript{2}SO\textsubscript{4} with NaOH, 60.0 mL of 0.020 M NaOH was needed to neutralize 15.0 mL of H\textsubscript{2}SO\textsubscript{4}. What is the molarity of the acid? (Be sure to write the neutralization reaction.)

7. If 10.0 mL of 0.300 M KOH are required to neutralize 30.0 mL of gastric juice (HCl), what is the molarity of the gastric juice?